

EP 34160 (4)

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property  
Organization  
International Bureau



(43) International Publication Date  
14 July 2005 (14.07.2005)

PCT

(10) International Publication Number  
WO 2005/063391 A1

(51) International Patent Classification<sup>7</sup>: B01J 35/00,  
23/38, 23/56

Kazuyuki. SUGA, Katsuo. NAKAMURA, Masanori.  
WAKAMATSU, Hironori. YASUDA, Hirofumi.

(21) International Application Number:  
PCT/JP2004/018333

(74) Agents: MIYOSHI, Hidekazu et al.; Toranomon Kotohira  
Tower, 2-8, Toranomon 1-chome, Minato-ku, Tokyo 105-  
0001 (JP).

(22) International Filing Date: 2 December 2004 (02.12.2004)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:  
2003-431009 25 December 2003 (25.12.2003) JP

(71) Applicant (for all designated States except US): NISSAN  
MOTOR CO., LTD. [JP/JP]; 2, Takara-cho, Kana-  
gawa-ku, Yokohama-shi, Kanagawa, 2210023 (JP).

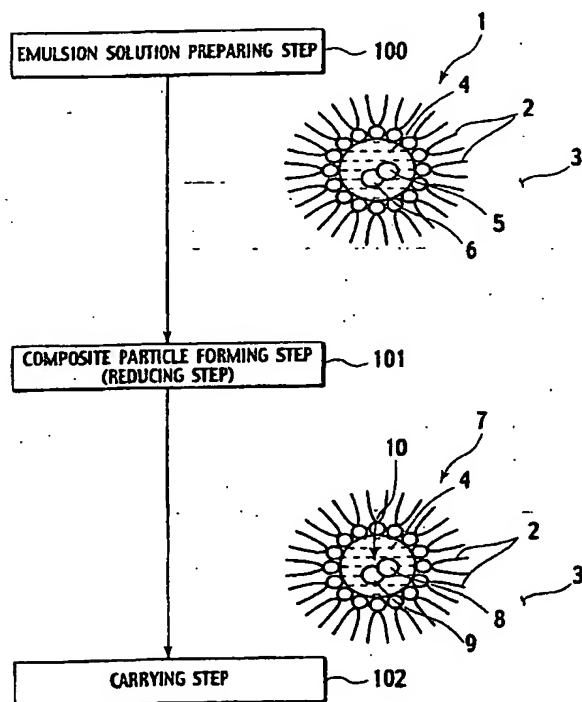
(81) Designated States (unless otherwise indicated, for every  
kind of national protection available): AE, AG, AL, AM,  
AT, AU, AZ, BA, BB, BG, BR, BW, BY, BZ, CA, CH, CN,  
CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, EG, ES, FI,  
GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, KE, KG,  
KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG,  
MK, MN, MW, MX, MZ, NA, NI, NO, NZ, OM, PG, PH,  
PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, SY, TJ, TM, TN,  
TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.

(72) Inventors; and  
(75) Inventors/Applicants (for US only): SHIRATORI,

(84) Designated States (unless otherwise indicated, for every  
kind of regional protection available): ARIPO (BW, GH,  
GM, KE, LS, MW, MZ, NA, SD, SL, SZ, TZ, UG, ZM,

[Continued on next page]

(54) Title: HEAT-RESISTIVE CATALYST AND PRODUCTION METHOD THEREOF



(57) Abstract: A high heat-resistant catalyst formed as a catalyst including a composite particle composed of a noble metal particle and a co-catalytic metal compound particle contacting, as a metal or as an oxide, with the noble metal particle, and a substrate carrying the noble metal particle and the co-catalytic metal compound particle, is produced by having a noble metal salt aqueous solution and a co-catalytic metal salt aqueous solution concurrently provided in a reverse micelle preparing reverse micellar solution containing a noble metal precursor and a co-catalytic metal precursor, and having a substrate carrying a composite particle comprising the noble metal precursor and the co-catalytic metal precursor concurrently reduced as a noble metal particle and a co-catalytic metal particle, respectively.

WO 2005/063391 A1



ZW), Eurasian (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM),  
European (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI,  
FR, GB, GR, HU, IE, IS, IT, LT, LU, MC, NL, PL, PT, RO,  
SE, SI, SK, TR), OAPI (BF, BJ, CF, CG, CI, CM, GA, GN,  
GQ, GW, ML, MR, NE, SN, TD, TG).

*For two-letter codes and other abbreviations, refer to the "Guid-  
ance Notes on Codes and Abbreviations" appearing at the begin-  
ning of each regular issue of the PCT Gazette.*

**Published:**

— with international search report

**DESCRIPTION**  
**HEAT-RESISTIVE CATALYST**  
**AND PRODUCTION METHOD THEREOF**

5

**TECHNICAL FIELD**

The present invention relates to a heat-resistive catalyst adaptive as an exhaust-gas purifying catalyser to be mounted on a vehicle, and to a production method thereof.

10

**BACKGROUND ART**

Along with a rising interest to the environment in recent years, the regulation to exhaust gases of automobile has been tightened with an increasing severity, involving improvements in engine system, accompanied by investigations for enhancement in performance of catalyst for purifying exhaust gases.

The exhaust-gas purifying catalyst has noble metal particles (e.g. of platinum (Pt), palladium (Pd)) held on surfaces of a carrying substrate (e.g. alumina ( $\text{Al}_2\text{O}_3$ )), for a conversion of harmful components (e.g. unburnt hydrocarbons (HC), carbon monoxide (CO)) contained in exhaust gases, by oxidization on noble metal particles, into harmless components (e.g. water, gas).

Noble metals to be active in catalysis are extremely expensive elements with an anxiety of resource exhaustion, which renders a large amount of use difficult.

Research and developments have been made for providing finer noble metal particles with extended surface areas to increase the contact area between exhaust gases and noble metal particles, allowing for noble metals to be reduced in amount

of use. Besides, attempts are made to have co-catalytic metal particles held on a carrying substrate.

A reversed micelle method (micro-emulsification) is known as an applicable technique to provide fine noble metal particles. The reversed micelle method first admix, in an organic solvent, a surfactant, and an aqueous solution containing e.g. noble metal element as a catalytically active component, for preparation of an emulsion having, in a reverse micelle formed in the organic solvent, the aqueous solution containing noble metal element. Noble metal is then precipitated, and reduced or rendered insoluble, so that reverse micelles have particles of noble metal formed therein, which are deposited.

Japanese Patent Application Laid-Open Publication No. 2000-42411 has disclosed a catalyst production method employing a reversed micelle method. First, particles of noble metal are formed by the reversed micelle method, and particles of co-catalytic metal (e.g. oxygen occluding metal particles). Then, by an impregnation method, a prescribed amount of particles of noble metal and a necessary amount of co-catalytic metal particles are held on a carrying substrate. Or, noble metal particles and co-catalytic metal particles are mixed together in a reduced or hydroxided state, to be held on a carrying substrate.

## 25 DISCLOSURE OF THE INVENTION

However, it is difficult for the method described to form a desirable composite of noble metal particles and co-catalytic metal particles, so that the contact area between noble metal particle and co-catalytic metal particle is reduced to be insufficient for co-catalytic metal particles to exhibit their

inherent performance (e.g. oxygen occluding ability). Further, the catalytic activity is reduced, with an amount of co-catalytic metal particles added with an amount of noble metal particles reduced.

5 In particular, upon exposure to a high temperature environment, the catalyst, which includes a substrate (e.g. metal oxide) carrying thereon noble metal particles and co-catalytic metal particles, has noble metal particles moved along substrate surfaces, sintering with the substrate, thus forming a complex  
10 oxide therebetween, resulting in a remarkable reduction in catalytic activity.

The present invention has been achieved as a solution to such points.

According to an aspect of the invention, a heat-resistive  
15 catalyst comprises a composite particle comprising a noble metal particle, and a co-catalytic metal compound particle contacting as a metal with the noble metal particle, and a substrate carrying the noble metal particle and the co-catalytic metal compound particle.

20 According to another aspect of the invention, a heat-resistive catalyst comprises a composite particle comprising a noble metal particle, and a co-catalytic metal compound particle contacting as an oxide with the noble metal particle, and a substrate carrying the noble metal particle and  
25 the co-catalytic metal compound particle.

According to another aspect of the invention, a production method of heat-resistive catalyst comprises having a noble metal salt aqueous solution and a co-catalytic metal salt aqueous solution concurrently provided in a reverse micelle, preparing  
30 reverse micellar solution containing a noble metal precursor

and a co-catalytic metal precursor, and having a substrate carrying a composite particle comprising the noble metal precursor and the co-catalytic metal precursor concurrently reduced as a noble metal particle and a co-catalytic metal particle, respectively.

The above and further objects, features, and advantages of the invention will more fully appear from the best mode for carrying out the invention, when the same is read in conjunction with the accompanying drawings.

10

#### BRIEF DESCRIPTION OF DRAWINGS

FIG. 1 is a schematic process chart of a method of producing a heat-resistive catalyst by a co-reduction according to an embodiment of the invention.

15

FIG. 2 is a process chart of a method of producing a heat-resistive catalyst by a co-reduction in an aluminium isopropoxide (Al-isoP) clathrate compound.

20

FIG. 3 is a process chart of a method of producing a heat-resistive catalyst by a co-reduction in an Al nitrate clathrate compound.

FIG. 4 is a process chart of a method of producing a catalyst in a Comparative Example 2 by a co-reduction in an aluminium isopropoxide (Al-isoP) clathrate compound.

#### 25 BEST MODE FOR CARRYING OUT THE INVENTION

There will be described below heat-resistive catalysts and production methods thereof according to embodiments as best modes for carrying out the present invention.

Description is first made of the composition of a high heat-resistant catalyst according to an embodiment of the

30

invention.

The heat-resistive catalyst has a complex of noble metal particles and co-catalytic metal compound particles, as a fine composite particle held on a carrying substrate, i.e., carried  
5 by or on the substrate.

In a molecular-level observation in light of an aspect of the embodiment, the composite particle is composed with a particle of a co-catalytic metal compound, e.g., transition metal compound contacting as a metal, i.e. in a metal state, with a  
10 noble metal particle.

The contact between a noble metal particle and a co-catalytic transition metal compound particle in a metal state gives rise to an increased tendency for electrons to move in between, promoting a spillover effect, causing the noble metal  
15 particle to have an enhanced catalytic activity.

The spillover effect has an increased influence, as the contact area between noble metal particle and transition metal compound particle extends. It therefore is desirable to make smaller the sizes or diameters of noble metal particles and  
20 transition metal compound particles, rendering greater the contact area between noble metal and transition metal compound particles.

The formation of a composite particle shortens distances between molecules as well as atoms of involved noble metal  
25 particles and transition metal compound particles, with additional contribution to the extension of contact area between noble metal and transition metal compound particles.

For a catalyst having an extended contact area between noble metal and transition metal compound particles, it is  
30 ensured in the form of an exhaust gas purifying catalyser mounted

on a vehicle, that reaction gases reach the transition metal, even in use within a stoichiometric range where the amount of reductant is equivalent to the oxygen amount of reaction gases.

The catalyst, which contains transition metal compound particles, enters a reduced state with an increased tendency to be active in catalysis, having an enhanced catalytic activity relative to a case employing noble metal particles alone.

In a molecular-level observation in light of another aspect of the embodiment, the composite particle is composed with a particle of a Co-catalytic metal compound, e.g., rare earth element compound or Zr-containing compound, contacting as an oxide, i.e. in an oxide state, with a noble metal particle.

The contact between a noble metal particle and a rare earth element compound particle in an oxide state promotes an oxygen absorbing/desorbing effect. For a catalyst containing rare earth element compound particles, in the form of a vehicle-mounted exhaust gas purifying catalyser, it is ensured, even in variations from the stoichiometric range of reaction atmosphere, where the amount of reductant is equivalent to the oxygen amount, to an oxygen-lean range or to an oxygen-rich range, that the oxygen absorbing/desorbing effect of rare earth element compound particles has a sufficient influence to keep noble metal particles in a metal state with an enhanced activity in catalysis, while the oxygen absorbing/desorbing effect has a significant contribution to controlling noble metal particles against variations of atmosphere, which otherwise might cause a deactivation of such noble metal particles due to a sintering or a transition to a solid solution in the substrate (porous oxide).

The oxygen absorbing/desorbing effect has an increased



influence, with an extended area by which rare earth element compound particles contact in an oxide state with noble metal particles.

5 It therefore is desirable to make smaller the sizes or diameters of noble metal particles and rare earth element compound particles, rendering greater the contact area between noble metal and rare earth element compound particles.

The formation of a composite particle having a Zr-containing particle contacting in an oxide state with a noble  
10 metal particle provides an excellent catalyst in anti-corrosion.

In the aspects of embodiment described, the co-catalytic transition metal compound particle may preferably be one of a simplex oxide, a complex oxide, a metal (of a 0 valence), and an alloy.

15 Metal compounds therefore may preferably contain one or more transition metal elements selected from among Fe, Co, Ni, Cu, Ti, and W, with a preference to (an) optimal metal(s) in respect of the use or kind of catalyst. Such (a) preferable metal(s) should hardly form a solid solution with an oxide  
20 substrate, allowing for an enhanced catalytic activity per unit mass of noble metal particles.

The co-catalytic metal compound particle may preferably contain a compound of a rare earth element (e.g. Ce, La) or an element (e.g. Zr) having an oxygen absorbing/desorbing effect.

25 The noble metal particle may preferably contain one or more noble metals selected from among Ru, Rh, Pd, Ag, Ir, Pt, and Au, with a preference to (an) optimal noble metal(s) in respect of the use or kind of catalyst.

Any noble metal particle may contain one or more kinds  
30 of noble metal. For example, in production of a catalyst, a Pt

salt and an Rh salt may be mixed in a reverse micelle, where they may be reduced to form a Pt-Rh composite particle, which may be carried on a surface of a metal oxide. Among such noble metals, a particular preference may be given to Pt, Pd, and Rh that are relatively high in catalytic activity.

Any composite particle may have one or more kinds of co-catalytic metal. For example, in production of a catalyst, a Co salt and a Ni salt may be admixed in a reverse micelle, where they may be reduced to form a composite particle having (a) noble metal particle(s) and (a) transition metal particle(s) containing a Co oxide and a Ni oxide, which composite particle may be carried on a surface of a metal oxide.

The carrying substrate may preferably be a porous oxide material composed of one or more oxides selected from among an alumina, a cerium oxide, a titanium oxide, a zirconia, and a silica.

Description is now made of a production method of heat-resistive catalyst according to an embodiment of the invention.

As illustrated in FIG. 1, the production method of heat-resistive catalyst includes an emulsion (reverse micellar solution) preparing process 100, a composite particle forming process 101, and a particle carrying process 102.

The emulsion preparing process 100 includes a step of mixing, in an organic solvent, a surfactant, a noble metal salt aqueous solution, and a co-catalytic metal salt aqueous solution, thereby preparing an emulsion with a multiplicity of dispersed reverse micelles each having noble metal salt aqueous solution and co-catalytic metal salt aqueous solution coexisting therein.

As shown in FIG. 1, a respective reverse micelle 1 is formed

spherical by surfactant molecules 2, with a diameter of approx. tens of nm. The reverse micelle 1 has an exterior of an oil phase 3, and an interior of a water phase 4. The water phase 4 has one or more noble metal precursors 5 (noble metal salt) and one  
5 or more co-catalytic metal precursors 6 (e.g. transition metal salt) coexisting therein, so that the noble metal and co-catalytic metal precursors 5 and 6 are uniformly mixed at a molecule level in the interior of reverse micelle 1.

The composite particle forming process 101 includes a  
10 (reduction) step of mixing a reductant in the emulsion, for a concurrent reduction of a respective noble metal precursor 5 (noble metal salt) and a respective co-catalytic metal precursor 6 (e.g. transition metal salt) in reverse micelle 1 to form a composite particle 10 having a corresponding noble metal particle  
15 8 and a corresponding catalytic metal compound particle 9. FIG. 1 illustrates a reverse micelle 7 after reduction, which has a composite particle 10 formed therein with a noble metal particle 8 and a co-catalytic metal compound particle 9.

The particle carrying process 102 includes a step of having  
20 composite particles held on a carrying substrate. This step may preferably be performed by one of first to third methods described below, as it is selective depending on use or kind of metals or metal salts in reverse micelle.

The first method employs a hydrolyzate of metal alkoxide  
25 for clathration. More specifically, to provide a precursor of the substrate (porous oxide), a metal alkoxide or a hydrolyzate of metal alkoxide is mixed in the emulsion, where it enters reverse micelles, whereby composite particles are mixed therewith to provide a resultant mixture.

30 Then, a solvent of the mixture is removed to provide dried

powder, which is fired to provide catalyst powder.

As the result, a heat-resistive catalyst is obtained, in which composite particles (of noble metal particle and co-catalytic metal compound particle) several nm to several tens  
5 nm in size are carried by surfaces of a substrate composed of (a) metal oxide(s).

The second method employs a precipitant for insolubilization of substrate salt. More specifically, a aqueous solution of a salt of a precursor of substrate (porous oxide)  
10 and a precursor salt of substrate (porous oxide) are precipitated or insolubilized as a hydroxide by a mixing of precipitant or insolubilizer, before a firing to have composite particles (noble metal particle and co-catalytic metal compound particle) carried by substrate surfaces.

15 The third method employs an impregnation for a carrying on substrate powder. More specifically, porous oxide powder is dispersed in a mixed solution, before a firing to have composite particles carried by surfaces of porous oxide (substrate).

Among the three methods described, the first method is  
20 most preferable, while the second method also is preferable. This is because, in the first method, alkoxide of porous oxide precursor is soon insolubilized by hydrolysis upon intrusion into reverse micelles, so that a manifest of insoluble alkoxide encloses composite particles (noble metal particle,  
25 co-catalytic metal compound particle), acting as a buffer for those particles. Catalysts produced by use of the first method can thus control noble metal particles against sintering, even under a high temperature condition.

In the production method of heat-resistive catalyst  
30 described, noble metal particles and co-catalytic metal compound

particles are prepared in advance as water-soluble salts. In other words, noble metal precursors (noble metal ions) and co-catalytic metal precursors (metal ions) are uniformly mixed at a molecule level, before their reduction to form composite particles about several nm to several tens nm in size or diameter, in which co-catalytic metal compound particles contact in a metal state with noble metal particles.

As a result, noble metal particles and co-catalytic metal particles are rendered smaller in size or diameter, having an increased contact area between noble metal particle and co-catalytic metal compound particle, allowing an enhanced activity of the latter. When carried by substrate surfaces, such composite particles can be controlled from aggregation.

The metal compound particles have different activated states depending on the element or working conditions. To achieve a metal compound state, conditions may be changed, for example of reactant kind, reaction temperature, reaction time, stirring strength, and stirring method.

Description is now made of materials to be used in the production method of heat-resistive catalyst.

The noble metal salt may be one of dinitro-diammine Pt(II) nitric acid-acidic aqueous solution, hexachloro Pt(IV) acidic solution, hexaammine Pt(IV) tetrachloride solution, Pd chloride aqueous solution, palladium nitrate aqueous solution, dinitro-diammine Pd dichloride solution, rhodium chloride solution, rhodium nitrate solution, ruthenium chloride solution, ruthenium nitrate solution, and hexachloro iridic acid aqueous solution, for example.

Organic solvent may be one of cyclohexane, methylcyclohexane, cycloheptane, heptanol, octanol, dodecyl

alcohol, cetyl alcohol, isooctane, n-heptane, n-hexane, n-decane, benzene, toluene, xylene, etc. There may be used a mixed solution of two or more of them. For example, reverse micelles containing noble metal particles and those containing  
5 co-catalytic metal compound particles may have different solutions prepared for use in the oil phase.

The surfactant may be one of polyoxyethylene nonylphenyl ether, magnesium laurate, zinc caprate, zinc myristate, sodium phenyl stearate, aluminum dicaprylate, tetra-isoamyl ammonium  
10 thiocyanate, n-octadecyl tri-n-butyl ammonium formate, n-amyl tri-n-butyl ammonium iodide, sodium bis-(2-ethylhexyl) succinate, sodium dinonyl naphthalene sulfonate, calcium cetyl sulfate, dodecyl amine oleate, dodecyl amine propionate, cetyltrimethylammonium bromide, stearyl trimethylammonium  
15 bromide, cetyltrimethylammonium chloride, stearyl trimethylammonium chloride, dodecyl trimethylammonium bromide, octadecyl trimethylammonium bromide, dodecyl trimethylammonium chloride, octadecyl trimethylammonium chloride, di-dodecyl dimethylammonium bromide, di-tetradecyl dimethylammonium  
20 bromide, di-dodecyl ditrimethylammonium chloride, di-tetradecyl dimethylammonium chloride, (2-octyloxy-1-octyloxymethyl)polyoxyethylene ethyl ether, etc. There may be used a mixed solution of two or more of them. For example, reverse micelles containing noble metal particles and  
25 those containing co-catalytic metal compound particles may have different solutions used for preparation of the surfactant.

The reductant may be one of hydrazine, sodium hydroborate, sodium thiosulfate, citric acid, sodium citrate, L-ascorbic acid, sodium borohydride, formic acid, formaldehyde, methanol,  
30 ethanol, ethylene, vitamin B, etc. There may be used a mixed

solution of two or more of them.

The precipitant may be ammonia water, tetramethylammonium hydroxide, or the like that is available to obtain hydroxides of associated noble metal and co-catalytic metal.

5           Examples

Examples of embodiment will be described with reference to FIG. 2 through FIG. 4.

Example 1

10           In Example 1, a catalyst powder was created by co-reduction in aluminium isopropoxide (Al-isoP) clathrate compound in FIG. 2.

15           Added to 66g of polyethylene glycol (5) mono-4-nonylphenyl ether as a surfactant was 1000ml of cyclohexane as solvent, thereby preparing a solution including 0.15mol% of surfactant, and the solution was then stirred. Meanwhile, added into 7.73ml of pure water were 0.37g of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt concentration was 8.46wt.%) and 0.26g of cobalt nitrate hexahydrate powder, and they were thereafter mixed and stirred. Thereafter, the prepared solutions were mixed, and  
20           stirred for about 2 hours, thereby subsequently obtaining reverse micellar solution including ions of Pt and Co (step 10).

25           Next, 0.12g of  $\text{NaBH}_4$  was added to the emulsion, followed by stirring for 2 hours, thereby obtaining reverse micellar solution including reduced composite particles (Pt-Co) (step 11).

30           Further, there was prepared a mixed solution by adding 20ml of cyclohexane to 4.0g of aluminium isopropoxide, and the prepared mixed solution was dropped into the emulsion including reduced Pt-Co composite particles, followed by stirring for about 2 hours. After stirring, reduced Pt-Co composite particles were

clathrated by Al hydroxide (step 12).

120ml of methanol was dropped into the emulsion including Pt-Co composite particles clathrated by Al hydroxide, thereby breaking reverse micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed by alcohol to remove excessive surfactant. This precipitate was dried at 100°C for 12 hours (step 13), and fired at 400°C in airflow (step 14), thereby obtaining catalyst powder in which 3wt.% of Pt and 5wt.% of Co were carried on every 1g of Al<sub>2</sub>O<sub>3</sub>.

Next, 50g of the catalyst powder obtained by repeating the above manipulations, 5g of boehmite, and 157g of 10% nitric-acid-containing aqueous solution were charged into an alumina-made porcelain pot, and shaken and ground together with alumina balls, thereby obtaining a catalyst slurry. Further, the catalyst slurry was loaded onto a honeycomb substrate made of cordierite (900 cell/2.5 mil, 0.06L) and an excessive slurry was removed by airflow, followed by drying at 120°C and firing thereafter at 400°C in airflow, thereby coating the catalyst to obtain a catalyst of Example 1.

#### Example 2

The same procedure as Example 1 was used except that 0.16g of hydrazine was added instead of NaBH<sub>4</sub> in the step 11 in Example 1, thereby creating catalyst powder of Example 2. Further, 500g of thus obtained catalyst powder was used and coated onto a honeycomb carrier like to the procedure of Example 1, thereby obtaining a catalyst of Example 2.

#### Example 3

In Example 3, a catalyst powder was created by co-reduction in Al nitrate clathrate compound in FIG. 3.



The same procedure as Example 1 was used up to the step 11, while using a nickel nitrate hexahydrate powder as the metal in the step 10 for Example 1. Here, an Al nitrate solution obtained by adding 7.36g of Al nitrate to 2ml of pure water, was added and mixed into a solution obtained by adding 225.7ml of cyclohexane to 14.9g of polyethylene glycol (5) mono-4-nonylphenylether, followed by stirring for about 2 hours, thereby preparing reverse micellar solution containing Al nitrate.

10 The reverse micellar solution including Pt-Ni composite particle and the reverse micellar solution Al nitrate were mixed, followed by stirring for about 2 hours, thereby obtaining reverse micellar solution in which Pt-Ni composite particle is mixed with Al nitrate (step 15).

15 Dropped into this emulsion was 70.5g of 25% ammonia water, whereby the Al nitrate was insolubilized as Al hydroxide, followed by further stirring for about 2 hours (step 16).

122.6ml of methanol was added to the prepared mixed solution to thereby break micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed by alcohol to remove excessive surfactant. Further, this precipitate was dried at 100°C for 12 hours (step 17), and then fired at 400°C in airflow (step 18), thereby obtaining catalyst powder in which 3wt.% of Pt and 5wt.% of Ni were carried on every 1g of Al<sub>2</sub>O<sub>3</sub>.

50g of thus obtained catalyst powder was used and carried on a honeycomb carrier by the same procedure as Example 1, thereby obtaining a catalyst of Example 3.

#### Example 4

30 In Example 4, a catalyst powder was created by a procedure

of impregnation into  $\text{Al}_2\text{O}_3$ .

The same procedure as the step 10 and step 11 of Example 1 was used, thereby obtaining reverse micellar solution including reduced Pt-CO composite particles.

5 Further, a mixed solution obtained by dispersing 1g of  $\gamma\text{-Al}_2\text{O}_3$  into 20ml of cyclohexane and the reverse micellar solution including reduced Pt-CO composite particles are mixed, for adsorption of reverse micelles to be carried on surfaces of  $\text{Al}_2\text{O}_3$ , followed by further stirring for 2 hours.

10 122.6ml of methanol was added into the prepared mixed solution to thereby breaking reverse micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed by alcohol to remove excessive surfactant. Further, this precipitate was  
15 dried at  $100^\circ\text{C}$  for 12 hours, and then fired at  $400^\circ\text{C}$  in airflow, thereby obtaining catalyst powder in which 3wt.% of Pt and 5wt.% of Co were carried on every 1g of  $\text{Al}_2\text{O}_3$ .

50g of catalyst powder obtained by repeating the above manipulations was used and carried on a honeycomb carrier by  
20 the same procedure as Example 1, thereby obtaining a catalyst of Example 4.

#### Example 5

In Example 5, a catalyst powder was created by co-reduction in Al nitrate clathrate compound in FIG. 3 to add Ce afterward.

25 There was prepared a solution including 0.15mol%/L of surfactant by adding 1,000ml of cyclohexane to 66g of polyethylene glycol(5) mono-4-nonylphenyl ether, followed by stirring. Meanwhile, added into 7.64ml of pure water were 0.44g of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt  
30 concentration: 8.46wt.%), and 0.46g of iron nitrate nonahydrate

powder, and then they were mixed and stirred. Thereafter, the solution including noble metal and transition metal was added into the prepared solution including surfactant, and they were stirred for about 2 hours, thereby subsequently obtaining reverse micellar solution including ions of Pt and Fe.

Next, added into the reverse micellar solution including Pt and Fe ions was 0.12g of  $\text{NaBH}_4$ , followed by stirring for 2 hours, thereby obtaining reverse micellar solution including reduced Pt-Fe composite particles.

Then, a solution obtained by mixing 7.36g of aluminum nitrate nonahydrate, 1.05g of cerium nitrate hexahydrate and 5.4ml of pure water was added to be mixed in a solution obtained by adding 237.1ml of cyclohexane to 15.7g of polyethylene glycol (5) mono-4-nonylphenyl ether, was, followed by stirring for about 2 hours.

The prepared emulsion including Pt-Fe composite particles and the prepared reverse micellar solution including Al nitrate and Ce nitrate were mixed, followed by stirring for about 2 hours, thereby subsequently obtaining reverse micellar solution including Pt-Fe composite particles and Al nitrate and Ce nitrate. Dropped into this solution was 71g of 25% ammonia water to thereby insolubilize Al nitrate and Ce nitrate, followed by stirring for about 2 hours.

122.6ml of methanol was added to the prepared mixed solution, thereby breaking the reverse micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed by alcohol to remove excessive surfactant. This precipitate was dried at  $100^\circ\text{C}$  for 12 hours, and then fired at  $400^\circ\text{C}$  in airflow, thereby obtaining catalyst powder.

50g of the catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier by the same procedure as Example 1, thereby obtaining a catalyst of Example 5.

5 Example 6

In Example 6, a catalyst powder was created by Pt-Co-Ce co-reduction in aluminium isopropoxide (Al-isoP) clathrate compound in FIG. 2.

There was prepared a solution including 0.15mol%/L of  
10 surfactant by adding 1,000ml of cyclohexane to 66g of polyethylene glycol(5)-mono-4-nonylphenyl ether, and this was stirred. Meanwhile, added into 7.56ml of pure water were 0.37g of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt concentration: 8.46wt.%), and cobalt nitrate hexahydrate powder  
15 and cerium nitrate, and they were mixed and then stirred. Thereafter, a solution including noble metal and transition metals was added into the prepared solution including surfactant, and they were stirred for about 2 hours, thereby subsequently preparing an emulsion with Pt-Co-Ce reverse micelles including  
20 ions of Pt, Co and Ce.

Next, added into the emulsion with Pt-Co-Ce reverse micelles was 0.12g of  $\text{NaBH}_4$ , followed by stirring for 2 hours, thereby obtaining an emulsion including Pt-Co-Ce composite particles.

25 Meanwhile, there were prepared a solution having 7.36g of aluminum nitrate nonahydrate mixed in 5.2ml of pure water, and a solution having 15.7g of polyethylene glycol(5) mono-4-nonylphenyl ether mixed in 226ml of cyclohexane, and then the prepared solutions were mixed together, followed by stirring  
30 for about 2 hours.

Added and mixed to the reverse micellar solution including Pt-Co-Ce composite particles was reverse micellar solution including Al nitrate, followed by stirring for about 2 hours, thereby obtaining reverse micellar solution having Pt-Co-Ce composite particles and Al nitrate mixed therein. 5  
Dropped into this emulsion was 71g of 25% ammonia water, thereby insolubilizing Al nitrate, followed by further stirring for about 2 hours.

Further, after preparation for mixing Pt-Co-Ce composite particles and insolubilized Al nitrate in reverse micelles, 10  
~~122.6ml of methanol was added to break the reverse micelles,~~  
followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed by alcohol to remove excessive surfactant. This precipitate was 15  
dried at 100°C for 12 hours, and then fired at 400°C in airflow, thereby obtaining catalyst powder containing Pt-Co-Ce composite particles carried on surfaces of Al<sub>2</sub>O<sub>3</sub>.

50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier by 20  
the same procedure as Example 1, thereby obtaining a catalyst of Example 6.

#### Example 7

In Example 7, a catalyst powder was created by Pt-Rh-Co co-reduction in aluminium isopropoxide (Al-isoP) clathrate 25  
compound in FIG. 2.

There was prepared a solution including 0.15mol%/L of surfactant by adding 1,000ml of cyclohexane to 66g of polyethylene glycol(5) mono-4-nonylphenyl ether, and this was stirred. Meanwhile, added into 7.56ml of pure water were 0.37g 30  
of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt

concentration: 8.46wt.%) and 0.103g of rhodium nitrate aqueous solution (Rh concentration: 5.83wt.%), and 0.26g of a cobalt nitrate hexahydrate powder, and they were mixed and then stirred.

Thereafter, the solution including the noble metals and  
5 the metal was added into the prepared solution including surfactant, and they were stirred for about 2 hours, thereby subsequently preparing reverse micellar solution including ions of Pt, Rh and Co.

Next, added into the emulsion was 0.12g of  $\text{NaBH}_4$ , followed  
10 by stirring for 2 hours, thereby obtaining reverse micellar solution including reduced Pt-Rh-Co composite particles.

Further, there were prepared a solution having 7.36g of aluminum nitrate nonahydrate mixed in 5.2ml of pure water, and a solution having 15.7g of polyethylene glycol (5)  
15 mono-4-nonylphenyl ether added to 226.0ml of cyclohexane, and then the prepared solutions were mixed together, followed by stirring for about 2 hours.

Added and mixed to the prepared reverse micellar solution including Pt-Rh-Co composite particles was the prepared reverse  
20 micellar solution including Al nitrate, followed by stirring for about 2 hours, thereby obtaining reverse micellar solution having Pt-Rh-Co composite particles and Al nitrate mixed therein. Dropped into this emulsion was 71g of 25% ammonia water to thereby insolubilize Al nitrate, followed by stirring for about 2 hours.

25 122.6ml of methanol was added to the reverse micellar solution having Pt-Rh-Co composite particles and insolubilized Al nitrate mixed therein, thereby breaking the reverse micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent. Thus obtained precipitate was washed  
30 by alcohol to remove excessive surfactant. This precipitate was

dried at 100°C for 12 hours, and then fired at 400°C in airflow, thereby obtaining catalyst powder containing Pt-Rh-Co composite particles carried on surfaces of Al<sub>2</sub>O<sub>3</sub>.

5 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier by the same procedure as Example 1, thereby obtaining a catalyst of Example 7.

#### Example 8

In Example 8, a catalyst powder was created by the same  
10 procedure as Example 1, except for adding iron nitrate  
nonahydrate instead of cobalt nitrate hexahydrate in the step  
10 of Example 1, so that the concentration of particles carried  
on Al<sub>2</sub>O<sub>3</sub> became 5wt.%, thereby obtaining catalyst powder. Further,  
50g of thus obtained catalyst powder was used and coated onto  
15 a honeycomb carrier, thereby obtaining a catalyst of Example  
8.

#### Example 9

In Example 9, a catalyst powder was created by the same  
procedure as Example 1, except for adding nickel nitrate  
20 hexahydrate instead of cobalt nitrate hexahydrate in the step  
10 of Example 1, so that the concentration of particles carried  
on Al<sub>2</sub>O<sub>3</sub> became 5wt.%, thereby obtaining catalyst powder. Further,  
50g of catalyst powder obtained by repeating the above  
manipulations was used and coated onto a honeycomb carrier,  
25 thereby obtaining a catalyst of Example 9.

#### Example 10

In Example 10, a catalyst powder was created by the same  
procedure as Example 1, except for adding a palladium nitrate  
aqueous solution instead of the dinitro-diamine Pt nitric  
30 acid-acidic aqueous solution and a lanthanum oxide nitrate

hexahydrate instead of the cobalt nitrate hexahydrate in the step 10 of Example 1, so that the concentrations of particles carried on  $\text{Al}_2\text{O}_3$  were 3wt.% and 5wt.%, respectively, thereby obtaining catalyst powder. Further, 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier, thereby obtaining a catalyst of Example 10.

#### Example 11

In Example 11, a catalyst powder was created by the same procedure as Example 1, except for adding a rhodium nitrate aqueous solution instead of the dinitro-diamine Pt nitric acid-acidic aqueous solution and a zirconium(IV) oxide nitrate hydrate instead of the cobalt nitrate hexahydrate in the step 10 of Example 1, so that the concentrations of particles carried on  $\text{Al}_2\text{O}_3$  were 3wt.% and 5wt.%, respectively, thereby obtaining catalyst powder. Further, 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier, thereby obtaining a catalyst of Example 11.

#### Example 12

In Example 12, a catalyst powder was created by the same procedure as Example 1, except for adding dinitro-diamine Pt nitric acid-acidic aqueous solution in the step 10 of Example 1, so that the concentration of particles carries on  $\text{Al}_2\text{O}_3$  became 3wt.%, thereby obtaining catalyst powder. Further, 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier, thereby obtaining a catalyst of Example 12.

#### Comparative Example 1

In Comparative Example 1, a catalyst including noble metal only was created in a similar procedure to Example 1.



There was prepared a solution including 0.15mol%/L of surfactant by adding 1,000ml of cyclohexane to 66g of polyethylene glycol (5) mono-4-nonylphenyl ether, and this was stirred. Meanwhile, added into 7.7ml of pure water was 0.37g of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt concentration: 8.46wt.%), and they were mixed and then stirred.

Thereafter, the solution including noble metal was added into the prepared solution including surfactant, and it was stirred for about 2 hours, thereby subsequently preparing reverse micellar solution including Pt ion.

Next, added into the reverse micellar solution including Pt ion was 0.024g of  $\text{NaBH}_4$ , followed by stirring for 2 hours, thereby obtaining reverse micellar solution including reduced Pt particles.

Further, there was prepared a mixed solution by adding 20ml of cyclohexane to 4.0g of aluminium isopropoxide, and the prepared mixed solution was dropped into the reverse micellar solution including reduced reduced Pt particles, followed by stirring for about 2 hours.

After stirring, Pt particles were clathrated by Al hydroxide in reverse micelles.

100ml of methanol was added to the reverse micellar solution including Pt particles clathrated by Al hydroxide, to break the reverse micelles, followed by stirring for about 2 hours, and filtration for separation from the solvent.

Thus obtained precipitate was washed by alcohol to remove excessive surfactant. This precipitate was dried at 100°C for 12 hours, and then fired at 400°C in airflow, thereby obtaining catalyst powder in which 3wt.% of Pt was carried on every 1g of  $\text{Al}_2\text{O}_3$ .

Further, 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier by the same procedure as Example 1, thereby obtaining the catalyst of Comparative Example 1.

5 Comparative Example 2

In Comparative Example 2, a catalyst powder was created by a co-reduction of Pt in aluminium isopropoxide (Al-isop) clathrate compound in FIG. 4, followed by a separate reduction for reducing Co.

10 As shown in FIG. 4, there was first prepared a solution including 0.15mol%/L of surfactant by adding cyclohexane to polyethylene glycol (5) mono-4-nonylphenyl ether, and this was stirred. Added and mixed to this solution were dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt concentration:  
15 8.46wt.%) and  $N_2H_4$ , followed by stirring for about 2 hours, thereby preparing reverse micellar solution including reduced Pt (step 19).

Next, there was prepared a solution including 0.15mol%/L of surfactant by adding cyclohexane to polyethylene glycol (5)  
20 mono-4-nonylphenyl ether like step 19, and this was stirred. Added and mixed into this solution were cerium nitrate and  $NaBH_4$ , followed by stirring for about 2 hours, thereby preparing reverse micellar solution including cerium hydroxide (step 20).

Thereafter, the emulsions prepared at step 19 and step  
25 20 were dropped into a cyclohexane mixed solution including aluminium isopropoxide, thereby clathrating Pt and Co by Al hydroxide, followed by stirring for about 2 hours (step 21).

100ml of methanol was added to the mixed solution prepared at step 21, to break the reverse micelles, followed by stirring  
30 for about 2 hours, and filtration for separation from the solvent.

Thus obtained precipitate was washed by alcohol to remove excessive surfactant (step 22). The precipitate was dried at 100°C for 12 hours (step 23), and then fired at 400°C in airflow (step 24), thereby obtaining catalyst powder in which 3wt.% of Pt and 5wt.% of Ce were carried on every 1g of  $\text{Al}_2\text{O}_3$ .

Further, 50g of catalyst powder obtained by repeating the above manipulations was used and coated onto a honeycomb carrier like to Example 1, thereby obtaining the catalyst of Comparative Example 2.

#### 10 Comparative Example 3

In Comparative Example 3, a catalyst powder was created by impregnation of Co followed by impregnation of Pt, using the impregnation method to  $\text{Al}_2\text{O}_3$ .

1g of  $\gamma\text{-Al}_2\text{O}_3$  was dispersed into a cobalt nitrate aqueous solution obtained by dissolving 0.26g of cobalt nitrate hexahydrate into 1.2ml of pure water, followed by stirring for about 1 hour. Thereafter, it was dried at a temperature of 150°C for 12 hours to remove pure water, and further fired at 400°C in airflow, thereby obtaining  $\text{Al}_2\text{O}_3$  having carried 5% of Co.

20 Next, the  $\text{Al}_2\text{O}_3$  powder having carried 5% of Co was dispersed in a solution obtained by adding 0.37g of dinitro-diamine Pt nitric acid-acidic aqueous solution (Pt concentration: 8.46wt.%) in 7.7ml of pure water, followed by stirring for about 1 hour, drying at 150°C for 12 hours, and firing at 400°C for 1 hour in airflow, thereby obtaining  $\text{Al}_2\text{O}_3$  powder having carried 3wt.% of Pt and 5wt.% of Co.

50g of thus obtained  $\text{Al}_2\text{O}_3$  catalyst powder having carried Pt and Co was used and coated onto a honeycomb carrier by the same procedure as Example 1, thereby obtaining the catalyst of Comparative Example 3.

The catalysts obtained in Examples 1 through 12 and Comparative Examples 1 through 3 were subjected to the following catalyst durability test, before evaluation of their catalytic performances.

5       The durability test employed a V-type 6-cylinder engine manufactured by Nissan Motor Co., Ltd., for which a lead-free gasoline was used as a fuel.

10       A jig adapted to adjust the flow rate of exhaust gases to a honeycomb carrier was employed for application to test-pieces in the durability test, which were sampled as exhaust-gas-purifying-catalysts from the catalysts obtained in Examples 1 through 12 and Comparative Examples 1 through 3.

15       Each sample had an adjusted catalyst position to set a catalyst inlet temperature to 700°C, and the engine was operated for 50 hours.

After the durability test, each sample was cut for an evaluation of catalytic performance to a catalyst capacity of 40 cc.

20       For the evaluation of catalytic performance, the composition of gas was conditioned to be stiochiometric between oxygen and reductant amounts, and a reaction gas of a composition shown in Table 1 was used.

25

30

Table 1

Reaction Gas Components	
Composition	Concentration
NO	1,000ppm
CO	0.60%
H <sub>2</sub>	0.20%
O <sub>2</sub>	0.60%
CO <sub>2</sub>	13.90%
C <sub>3</sub> H <sub>8</sub>	1,665ppm
H <sub>2</sub> O	10
N <sub>2</sub>	balance

The performance evaluation for each catalyst having a catalyst capacity of 40cc was conducted at a flow rate of reaction gas of 40L/min and a reaction gas temperature of 350°C.

There was calculated an NOx purifying ratio (%) from a ratio based on a catalyst inlet NOx concentration and an outlet NOx concentration upon stabilization of these concentrations, and evaluation results are shown in Table 2 on the overleaf.

As shown in Table 2, there was exhibited the highest value of 58% for an NOx purifying ratio of Example 7 in which the catalyst powder was created by Pt-Rh-Co co-reduction in aluminium isopropoxide clathration, and the addition of Rh into the catalyst made it to have a catalytic activity higher than a situation of preparation by Pt only.

Although Comparative Example 1 carried Pt only and had an NOx purifying ratio of about 48%, the catalyst of Example 1 produced by the same procedure as Comparative Example 1 had an NOx purifying ratio increased to 54% by addition of Co as co-catalytic component into micelles in addition to Pt, thereby showing that its catalytic activity was enhanced as compared

with Comparative Example 1. Further, in Example 12, although the amount of used Pt was so decreased by setting the carried concentration of Pt to be 0.50%, it was possible to obtain an NOx purifying ratio of 49% by carrying a co-catalytic component on a substrate by setting the carried concentration of Co to be 5.0%.

As a result, it was shown that catalytic activities are enhanced by carrying co-catalytic components on substrates even when carried concentrations of noble metals are decreased and thus usage amounts of the noble metals are decreased.

While preferred embodiments of the present invention have been described using specific terms, such description is for illustrative purposes, and it is to be understood that changes and variations may be made without departing from the spirit or scope of the following claims.

Table 2

	Noble metal		Metal element (cocat. comp.)		Reducing agent	Metal element to be added in post-step		Precursor of porous oxide		Insolubilizing agent for metal (substrate)	Coating amount onto honeycomb (g/L)	$\eta$ Nox@350°C after endurance at 700°Cx30Hr
	Kind	Carried conc. (%)	Kind	Carried conc. (%)		Kind	Carried conc. (%)	Kind	Precursor			
Example	1	Pt	3.00	Co	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	54
	2	Pt	3.00	Co	5.0	N <sub>2</sub> H <sub>2</sub>	.	Al	Al isopropoxide	.	100	56
	3	Pt	3.00	Ni	5.0	NaBH <sub>4</sub>	.	Al	Al nitrate	ammo. water	100	52
	4	Pt	3.00	Co	5.0	NaBH <sub>4</sub>	.	Al	None (impreg. into Al <sub>2</sub> O <sub>3</sub> )	.	100	52
	5	Pt	3.00	Fe	5.0	NaBH <sub>4</sub>	Ce	Al	Al nitrate	ammo. water	100	51
	6	Pt	3.00	Co	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	51
				Ce	10.0							
	7	Pt	3.00	Zr	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	58
		Rh	0.60									
	8	Pt	3.00	Fe	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	51
	9	Pt	3.00	Ni	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	51
	10	Pd	3.00	La	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	50
Comparative Example	11	Rh	3.00	Zr	3.5	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	50
	12	Pt	0.50	Co	5.0	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	49
	1	Pt	3.00	.	.	NaBH <sub>4</sub>	.	Al	Al isopropoxide	.	100	48
	2	Pt	3.00	.	.	N <sub>2</sub> H <sub>2</sub>	Ce	Al	None (impreg. into Al <sub>2</sub> O <sub>3</sub> )	.	100	49
	3	Pt	3.00	.	.	.	Co (impreg.)	Al	None (impreg. into Al <sub>2</sub> O <sub>3</sub> )	.	100	50

**INDUSTRIAL APPLICABILITY**

According to the present invention, there is provided a heat-resistive catalyst in which composite particles containing noble metal and co-catalytic metallic compound are carried on a substrate, whereby a co-catalytic effect of metallic compound is kept active even with a reduced amount of noble metal, allowing for a low-cost catalyst relatively free of deterioration in catalytic activity.

Further, according to the present invention, there is provided a production method of heat-resistive catalyst employing a reversed micelle method of having noble metal salt and co-catalytic metal salt co-existing in a reverse micelle for formation of composite particle, whereby a co-catalytic effect of metallic compound is promoted, allowing for a heat-resistive catalyst high of catalytic activity and low of cost.



## CLAIMS

1. A heat-resistive catalyst comprising:

5       a composite particle comprising a noble metal particle,  
and a co-catalytic metal compound particle contacting as a metal  
with the noble metal particle; and

      a substrate carrying the noble metal particle and the  
co-catalytic metal compound particle.

10

2. A heat-resistive catalyst comprising:

      a composite particle comprising a noble metal particle,  
and a co-catalytic metal compound particle contacting as an oxide  
with the noble metal particle; and

15       a substrate carrying the noble metal particle and the  
co-catalytic metal compound particle.

3. The heat-resistive catalyst as claimed in claim 1, wherein  
the co-catalytic metal compound particle comprises a transition  
20 metal compound.

4. The heat-resistive catalyst as claimed in claim 2, wherein  
the co-catalytic metal compound particle comprises one of a rare  
earth element compound and a compound containing Zr.

25

5. The heat-resistive catalyst as claimed in claim 1 or 2, wherein  
the substrate comprises a porous oxide having a surface carrying  
the composite particle.

30 6. The heat-resistive catalyst as claimed in claim 1 or 2, wherein

the substrate comprises at least one porous oxide selected from among an alumina, a cerium oxide, a titanium oxide, a zirconia, and a silica.

5 7. The heat-resistive catalyst as claimed in claim 1 or 2, wherein the noble metal particle comprises at least one metal selected from among Ru, Rh, Pd, Ag, Ir, Pt, and Au.

10 8. The heat-resistive catalyst as claimed in claim 1 or 2, wherein the co-catalytic metal compound particle comprises a transition metal compound containing at least one transition metal selected from among Fe, Co, Ni, Cu, Ti, and W.

15 9. A production method of heat-resistive catalyst, comprising:  
having a noble metal salt aqueous solution and a co-catalytic metal salt aqueous solution concurrently provided in a reverse micelle,

preparing reverse micellar solution containing a noble metal precursor and a co-catalytic metal precursor; and  
20 having a substrate carrying a composite particle comprising the noble metal precursor and the co-catalytic metal precursor concurrently reduced as a noble metal particle and a co-catalytic metal particle, respectively.

25 10. The production method of heat-resistive catalyst as claimed in claim 9, comprising providing a reductant to the emulsion, concurrently reducing the noble metal precursor and the co-catalytic metal precursor in the reverse micelle, forming the composite particle.

11. The production method of heat-resistive catalyst as claimed in claim 9, comprising:

mixing, in the reverse micelle, a hydrolyzate of alkoxide as a precursor of a porous oxide forming the substrate, having  
5 a mixture; and

firing the mixture, before carrying the composite particle by a surface of the porous oxide.

12. The production method of heat-resistive catalyst as claimed

10 in claim 9, comprising mixing, in the reverse micelle, an aqueous solution of a precursor salt of a porous oxide forming the substrate and a precipitating agent or an insolubilizing agent for precipitating or insolubilizing the precursor salt of the porous oxide as a hydroxide, before a firing to carry  
15 the composite particle by a surface of the porous oxide.

13. The production method of heat-resistive catalyst as claimed in claim 9, comprising dispersing, in the emulsion, powder of a porous oxide forming the substrate, before a firing to carry  
20 the composite particle by a surface of the porous oxide.

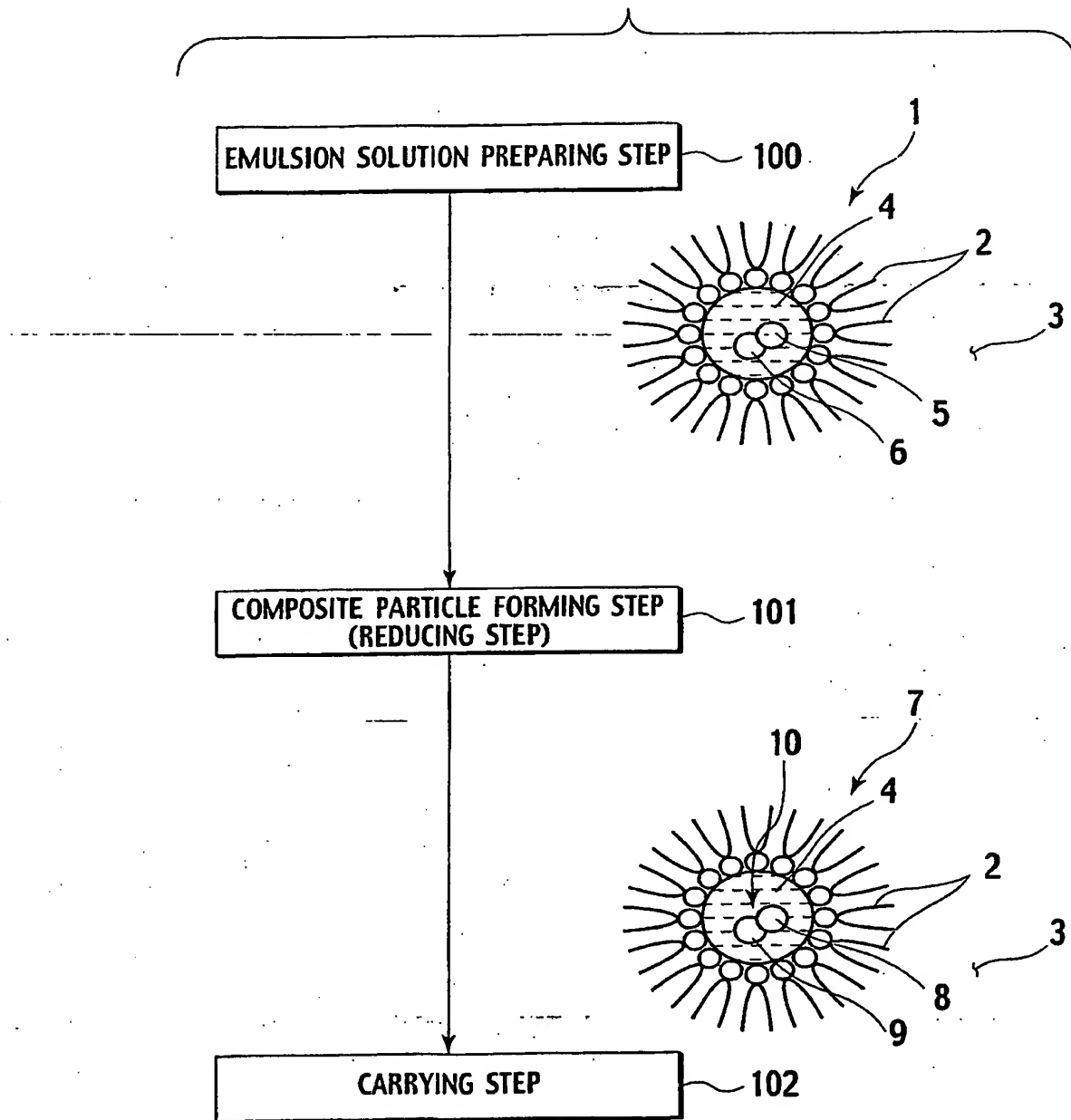
14. The production method of heat-resistive catalyst as claimed in claim 9, wherein the noble metal salt aqueous solution comprises a metal salt aqueous solution of at least one metal  
25 selected from among Ru, Rh, Pd, Ag, Ir, Pt, and Au.

15. The production method of heat-resistive catalyst as claimed in claim 9, wherein the co-catalytic metal salt aqueous solution comprises a metal salt aqueous solution of at least one metal  
30 selected from among Fe, Co, Ni, Cu, Ce, Zr, La, Ti and W.

16. The production method of heat-resistive catalyst as claimed  
in claim 9, wherein the substrate comprises a porous oxide  
containing at least one oxide selected from among an alumina,  
5 a cerium oxide, a titanium oxide, a zirconia, and a silica.

1/3

FIG.1



2/3

FIG.2

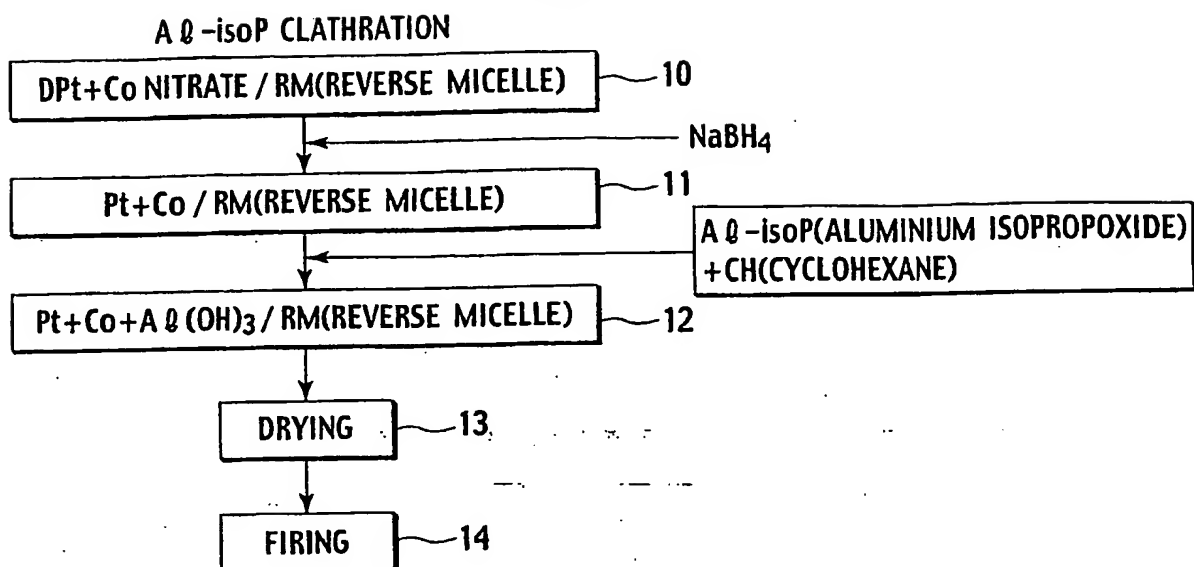
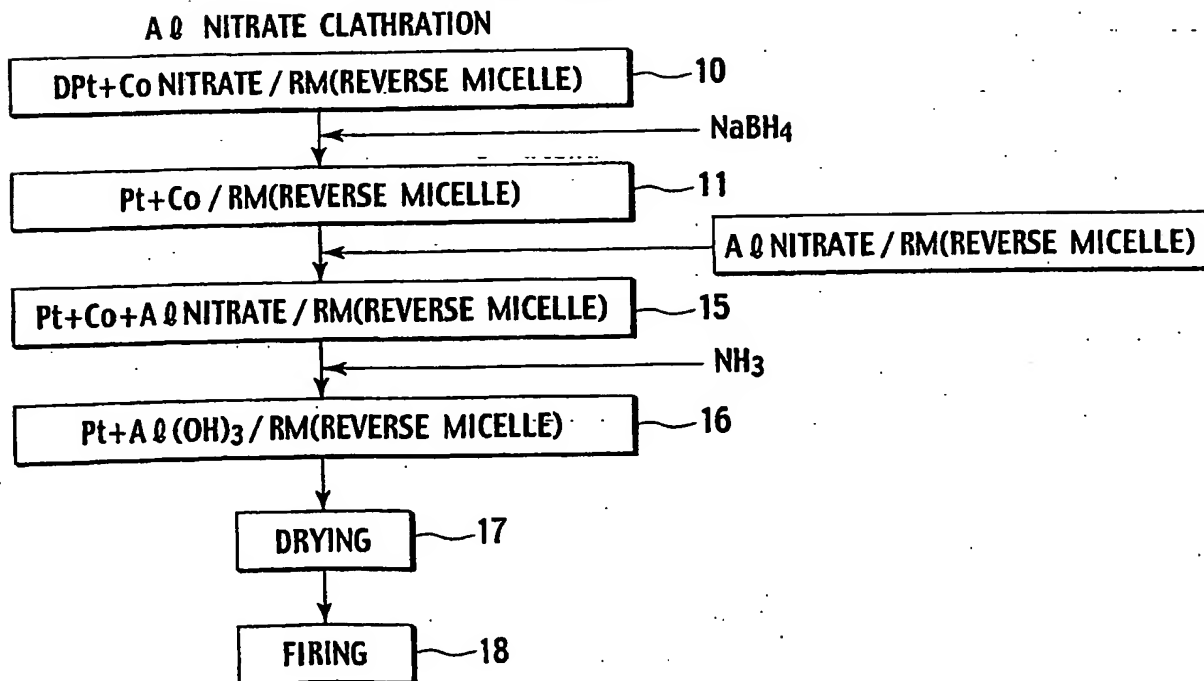
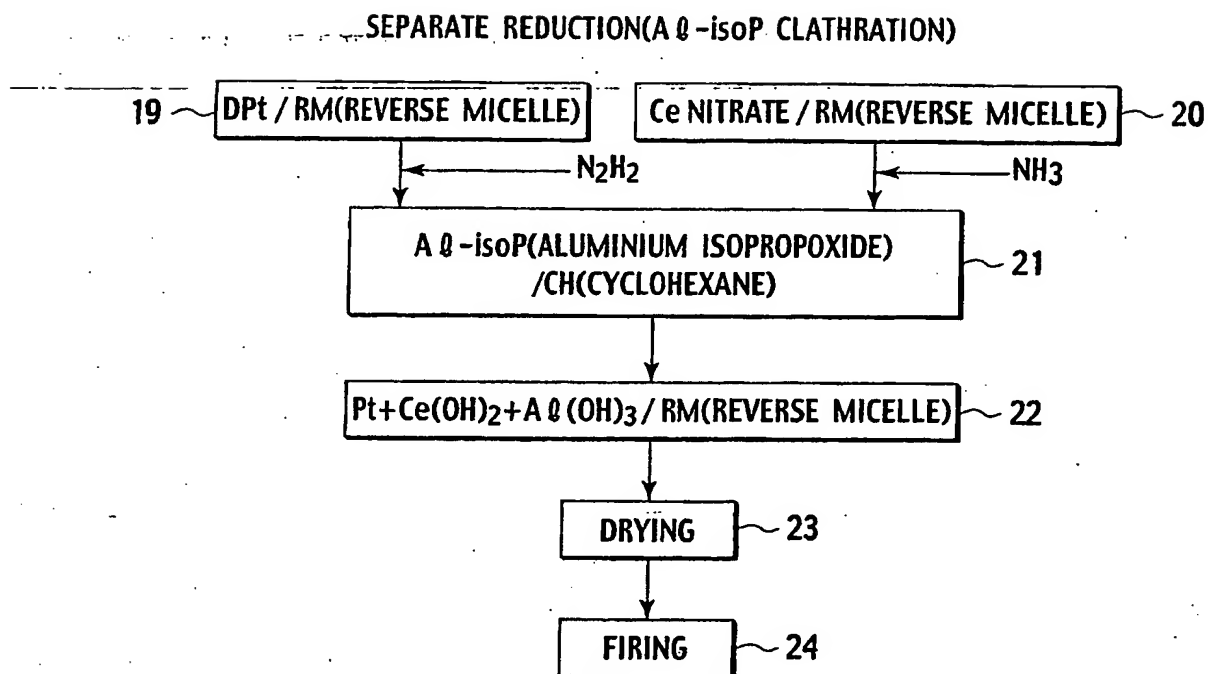


FIG.3



3/3

FIG.4



# INTERNATIONAL SEARCH REPORT

International Application No  
/JP2004/018333

A. CLASSIFICATION OF SUBJECT MATTER  
IPC 7 B01J35/00 B01J23/38 B01J23/56

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)  
IPC 7 B01J B01D H01M

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data, COMPENDEX, PAJ

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	PATENT ABSTRACTS OF JAPAN vol. 2000, no. 05, 14 September 2000 (2000-09-14) -& JP 2000 042411 A (TOYOTA MOTOR CORP; WAKABAYASHI KATSUHIKO), 15 February 2000 (2000-02-15) cited in the application paragraph '0025! - paragraph '0032!	1-16
X	EP 0 466 984 A (UOP) 22 January 1992 (1992-01-22) column 7, line 30 - column 8, line 7 column 9, line 7 - line 14	1
X	US 6 083 467 A (TAKESHIMA ET AL) 4 July 2000 (2000-07-04) column 8, line 55 - column 9, line 15; figures 4,5,7; example 5 ----- -/-	1,9

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

### \* Special categories of cited documents:

- \*A\* document defining the general state of the art which is not considered to be of particular relevance
- \*E\* earlier document but published on or after the international filing date
- \*L\* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- \*O\* document referring to an oral disclosure, use, exhibition or other means
- \*P\* document published prior to the international filing date but later than the priority date claimed

- \*T\* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- \*X\* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- \*Y\* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- \*Z\* document member of the same patent family

Date of the actual completion of the international search

15 February 2005

Date of mailing of the international search report

28/02/2005

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,  
Fax (+31-70) 340-3016

Authorized officer

Veefkind, V



# INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

/JP2004/018333

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
JP 2000042411	A	15-02-2000	NONE	
EP 0466984	A	22-01-1992	US 4945116 A EP 0466984 A1 US 5059574 A ZA 9005574 A	31-07-1990 22-01-1992 22-10-1991 24-04-1991
US 6083467	A	04-07-2000	JP 3466856 B2 JP 10216517 A DE 19804587 A1	17-11-2003 18-08-1998 06-08-1998
JP 7246343	A	26-09-1995	JP 3590854 B2	17-11-2004
EP 1494304	A	05-01-2005	EP 1494304 A1 US 2005003959 A1	05-01-2005 06-01-2005
EP 0931590	A	28-07-1999	JP 11207190 A JP 11276896 A DE 69919917 D1 EP 0931590 A1 US 6221804 B1	03-08-1999 12-10-1999 14-10-2004 28-07-1999 24-04-2001

Form PCT/ISA/210 (patent family annex) (January 2004)

BNSDOCID: <WO\_2005063391A1\_L>

## INTERNATIONAL SEARCH REPORT

International Application No  
/JP2004/018333

## C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	DATABASE WPI Section Ch, Week 199547 Derwent Publications Ltd., London, GB; Class A97, AN 1995-363054 XP002317375 -& JP 07 246343 A (WAKABAYASHI K) 26 September 1995 (1995-09-26) paragraphs '0015!', '0019!' - '0022!', '0025!', '0027!'	1-16
E,L	EP 1 494 304 A (NISSAN MOTOR CO., LTD) 5 January 2005 (2005-01-05)  (L: priority) claim 20; example 3	1-3, 7-10,14, 15
X	EP 0 931 590 A (MAZDA MOTOR CORPORATION) 28 July 1999 (1999-07-28) paragraph '0011!'	1-8
A	KIM W-Y ET AL: "Methanol synthesis from syngas over supported palladium catalysts prepared using water-in-oil microemulsion" APPLIED CATALYSIS A: GENERAL, ELSEVIER SCIENCE, AMSTERDAM, NL, vol. 169, no. 1, 17 April 1998 (1998-04-17), pages 157-164, XP004271387 ISSN: 0926-860X Paragraph 2.1 "Catalyst preparation"figure 1	1-16
A	MARTINEZ-ARIAS ET AL.: "Study of the lean NOx reduction with C3H6 in the presence of water over silver/alumina catalysts prepared from inverse microemulsions" APPLIED CATALYSIS B ENVIRONMENTAL, vol. 28, 2000, pages 29-41, XP002317374 the whole document	13

Form PCT/ISA/210 (continuation of second sheet) (January 2004)

page 2 of 2